

Acids and Alkalis

Glossary



A

acid:	An acid has a pH value of less than 7; a strong acid will turn universal indicator red.
alkali:	An alkali has a pH value of more than 7; a strong alkali will turn universal indicator blue.

B

base:	A substance that will neutralise an acid but does not dissolve in water.
burette:	A glass tube with a tap at one end, so that a known volume of liquid can be measured.

C

citric acid:	The acid that is commonly found in lemons.
concentrated:	A substance is concentrated if it has a high proportion of that substance compared to other substances in the solution.
corrosive:	This may dissolve or burn.

D

dilute:	A solution is often diluted by the addition of water.

E

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F

G

H

hazard symbols:

These are symbols designed to warn of potential dangers.

hydrochloric acid:

A type of acid often used in school science experiments. Its formula is HCl.

I

indicator:

A substance that changes colour to show whether a solution is acidic or alkaline.

J**K**

L

litmus paper:

There are two types of litmus paper, blue and red. They act as indicators.

M

N

neutral:

A solution that has a pH value of 7. Water is a good example.

neutralisation:

Adding an alkali to an acid to produce a neutral solution.

O

P

pH scale:

Used to determine whether a solution is acidic or alkaline (see illustration below).



Q**R****S**

sodium hydroxide:

An alkali often used in school science experiments. Its formula is NaOH.

sulphuric acid:

An acid often used in school science experiments. Its formula is H₂SO₄.

T**U**

universal indicator:

An indicator that changes colour to show the pH of a solution.

V

W

X

y

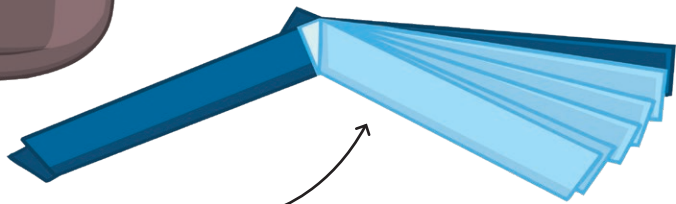
Z



strong, dilute
acids symbol



concentrated acids symbol



litmus paper

